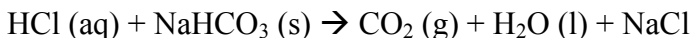


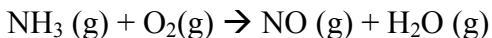
STOICHIOMETRY PRACTICE PROBLEMS

Make sure equations are balanced before working!!!!!!

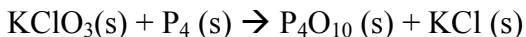
1. What mass of sodium chloride can be formed by reacting 5.00g of sodium bicarbonate with hydrochloric acid?



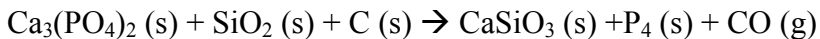
2. Ammonia reacts with oxygen to form nitrogen monoxide and water. What mass of nitrogen monoxide can be formed from 15.00 g of oxygen reacting with excess ammonia?



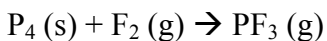
3. What mass of both P₄O₁₀ and KCl can be formed by reacting 52.9 g of potassium chlorate with excess phosphorous?



4. Phosphorous can be prepared from calcium phosphate by the following reaction:



5. What mass of F₂ is needed to form 120.0 g of PF₃?



6. If 100 g of P₄ reacts with 75 g F₂ in the above reaction, what is the limiting reagent? How many grams of PF₃ are then formed?

7. Bornite (Cu_3FeS_3) is a copper ore used in the production of copper. What masses of Cu, FeO, and SO_2 are formed by the reaction of 55.0 g of bornite with oxygen?



8. If 55.0 g bornite reacts with 10 g oxygen, what is the limiting reagent?

9. From your answer to #8 above, how many grams of Cu, FeO, and SO_2 are formed?

10. What mass of hydrogen cyanide can be formed by reacting 5.00×10^3 g of CH_4 with ammonia and oxygen?